**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 6 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 6 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. The atomic radius from left to right in the period: A) Decrease B)Increase C) Remain same D)None
2. Ionic radius appeared to be an property: A) Multiplicative B) Associative C) Additive D) None
3. Ionization energy is measure of stability of isolated form: A) Quantitative B) Qualitative C) Both D) None
4. The Ionization energy of the group show abnormal trend: A) 3, 6 B) 2,5,8 C) 3,8 D) 6,8
5. The highest electro negativity in a periodic table is of: A) Cl B) Cs C) H D) F
6. The electro negativity of chlorine is: A) -341 B) +341 C)-349 D)+349
7. NaCl has ---% Ionic Character: A) 72 % B) 92 % C) 100 % D) 0 %
8. The bond angle betweeb two H-S Bond is: A) 102 B) 105 C) 92 D) 109.5
9. The geometry of NF3 is: A) Bent B) Trigonalpyramidal C)Planer D) Linear
10. Number of bond in N2 is: A) 1 S and 2 p B) 1s and 3 p C) 2 S and 2p D)None

**Q # 2: Short Question 10 \* 2= 20**

1. Draw molecular orbital diagram of Helium (He).
2. Define bond order.
3. What is Ionic Bond?
4. Define Hybridization.
5. Define covalent bond according to quantum mechanical approach.
6. Write postulate of VSEPR theory.
7. What is co-ordinate covalent bond?
8. Define electro negativity.
9. Define electron affinity with example.
10. Why the radius of atom cannot determined with precision.

**Q # 3 Long Questions 2\* 5 =10**

1. Explain sp Hybridization with one example.
2. Explain ionization energy along with example and trend.