**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 11 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 11 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. Instantaneous and average rate become equal when time internal approaches to: A) 1 B) 0 C) Maximum D) None
2. 2N2O5 2N2O4 + O2 order is? A) 1 B) 2 C) 3 D) 4
3. Photochemical reaction are usually: A) 1st order B) 2nd order C) 3rd order D) zero
4. Half life period of N2O5 at 45 C is: A) 48 mint B) 24 mint C) 36 mint D) 100 mint
5. For endothermic reaction the product are at energy level reactant: A) Lower B) Higher C) Both D) None
6. The unit of slop in Arrhenius equation is: A) K-1 B) K0 C) K D) None
7. The rate of reaction as the reaction proceeds: A) Decrease B) Increase C) Same D) None
8. Arrhenius equation can be used find out reaction: A) Rate B) Order C) both D) none
9. Collision theory explain of reaction: A) Rate B) Order C) both D) none
10. Unit of rate of reaction is: A) mol-1dm-3s B) moldm-3s-1 C) both D) None

**Q # 2: Short Question 10 \* 2= 20**

1. Define rate of reaction.
2. What is instantaneous rate and average rate?
3. Define order of reaction.
4. What do you know about half life period?
5. What is rate limiting step?
6. Write methods of large excess.
7. Order of reaction is an experimental determined quantity why?
8. What is reaction intermediate?
9. What is zero reaction?
10. Write method to find order of reaction.

**Q # 3 Long Questions 2\* 5 =10**

1. Define and explain energy of activation.
2. Explain Arrhenius equation.